

General Chemistry I: Course Syllabus
Windward Community College – Spring 2014

Course Number: CHEM 161 (CRN 60132; 3 credits)

Class Meeting Days and Times: MW 11:30 am - 12:45 pm ('Imiloa 111)

Instructor: Dr. Christopher Guay

Email: cguay@hawaii.edu

Course website: <http://laulima.hawaii.edu> (use UH email account login and password)

Office Hours: MW 1:00-1:50 pm; T 11:30 am-12:20 pm; and R 12:00-12:50 pm ('Imiloa 136)

WINDWARD COMMUNITY COLLEGE MISSION STATEMENT

Windward Community College offers innovative programs in the arts and sciences and opportunities to gain knowledge and understanding of Hawai'i and its unique heritage. With a special commitment to support the access and educational needs of Native Hawaiians, we provide O'ahu's Ko'olau region and beyond with liberal arts, career and lifelong learning in a supportive and challenging environment – inspiring students to excellence.

CATALOG DESCRIPTION OF THE COURSE

Basic principles of inorganic chemistry with emphasis on problem solving. First course of a two-course sequence designed to meet the one-year General Chemistry requirement for pre-med, science and engineering majors. Topics include chemical calculations, electronic structure, chemical bonding, states of matter and solutions. (3 hrs. lecture)

Prerequisites: A grade of 'C' or better in Math 103, or placement in Math 135 or instructor's consent.

Co-requisite: Concurrent registration in Chem 161L.

Recommended Preparation: Student should have taken high school chemistry, Chem 100 or Chem 151.

WCC: DP

STUDENT LEARNING OUTCOMES

1. Use the mole concept in solving stoichiometry problems involving solids, liquids, gases and solutions.
2. Balance chemical equations, classify reactions, identify and analyze the role of the chemicals involved in chemical reactions.
3. Predict the behavior of gases while undergoing changes in volume, pressure, temperature and quantity.
4. Manipulate thermochemical equations and calculate the amount of energy involved in chemical reactions.
5. Predict physical and chemical properties of elements based on electronic structure and location in the Periodic Table.
6. Predict physical and chemical properties of compounds based on chemical bonding, geometry and intermolecular interactions.

COURSE TASKS

- Daily attendance
- Homework (online and paper-based)
- In-class quizzes
- Three midterm exams
- Final exam (ACS National Standardized Exam for General Chemistry I)
- At least eight supplemental instruction (SI) sessions

REQUIRED COURSE MATERIALS

- Chemistry in a Flash: College Chemistry by Ketan Trivedi. This is an electronic text (including interactive problem solving modules, animations, etc.) that is contained on a flash drive. Versions for Mac or PC are available. The bookstore may not have these available yet, but you can order them online directly at: <http://trivedichemistry.com/index.php/college-chemistry>
- Access to Mastering Chemistry for online homework and tutorials (purchase access key online or from WCC Bookstore). Go to www.masteringchemistry.com and register/login. Join our course using the course ID code **MCGUAY12480**.
- **IMPORTANT!** When you are prompted to select a text for the course, choose “**Tro, Chemistry: A Molecular Approach, 3e**”. You do not need to pay extra to purchase access to this e-text. You just need to select this text so that you will have access to the correct set of homework problems.
- You will also need a scientific calculator and Internet access.

GRADING

1. Grades will be based on the following categories:

- i. Attendance & Quizzes
- ii. Homework
- iii. Midterm Exam 1
- iv. Midterm Exam 2
- v. Midterm Exam 3
- vi. Final Exam (counts double – *i.e.*, counts as two categories)

Your percentage score in each category will be determined, and the category with the lowest score will be dropped. An average percentage score for the remaining six categories will be calculated and used to assign your grade for the course as follows:

- A: 100 - 90.0 %
- B: 89.9 - 80.0 %
- C: 79.9 – 70.0 %
- D: 69.0 – 60.0 %
- F: below 60 %

Curving might be employed if deemed necessary.

Grades of I, W, CR, NC are described in the current college catalog. The last day to change from letter grading (A-F) to CR/NC option is on November 4, 2013 – this must be discussed previously with the instructor.

2. **Attendance** will be checked during every class period. You are **required to attend every class** – you will lose points for missing class. **Supplemental Instruction (SI)** sessions will be offered after every class. You are required to attend **at least 8** of these sessions. You will lose points for not attending the minimum number of SI sessions, and you can earn extra credit points for attending more than the minimum number of SI sessions.
3. **Quizzes:** A short in-class **quiz** (1-3 questions) will be given during every lecture period.
4. **Homework assignments** (online and paper-based) will be due every class period.

Online homework: Online homework assignments will be given through the Mastering Chemistry website. Go to www.masteringchemistry.com and register/login. Join our course using the course ID code **MCGUAY12480**. Online homework assignments will be due every Monday (even if it's a holiday). Late homework submissions will be penalized 10% per day.

Written homework: Written homework problems will be assigned from the course DVD-ROM. The assignments must be done by hand (*i.e.*, pencil and paper) – be sure to show your work! Written homework assignments are due every Wednesday at the beginning of the class period. **Late homework will not be accepted for written homework assignments.** If you are not coming to

class on a day when a written homework assignment is due, you can still scan the assignment and email it to me – but it must be received before the start of the class period to receive credit.

5. There will be **three midterm exams**, each of which will cover approximately one-third of the course. Each exam will last for 75 minutes. All exams will be closed book.
6. The **final exam** will be the American Chemical Society's national, standardized exam for First-Term General Chemistry. This exam will consist of 70 multiple-choice questions and **cover all topics** presented in the course (*i.e.*, cumulative). You will be given 2 hours to complete the exam. The final exam will also be closed book.

HOW TO STUDY FOR THIS COURSE

Nothing is more important to your academic success than strong study skills. On average, you should spend about **seven hours per week** outside the classroom to study for this course.

1. Prepare for each class by familiarizing yourself with the **lecture slides**, which will be posted beforehand on the course website.
2. Take **notes** during the lecture, but don't focus too much energy on trying to write down every single thing (remember, you can download and print out the lecture slides). Stay engaged and participate in class, ask questions, etc. Bring your **calculator** to class so you can work through sample problems (you will also need it for the in-class quizzes and activities).
3. **Review** your notes soon after class. Attend the **supplemental instruction** session held after every lecture in the classroom. This is a good place to get tips on how to solve your homework and review for quizzes and exams.
4. Watch your **DVD** and do all of the **interactive problems**. Work through the **online tutorials** on the Mastering Chemistry site. Summarize the ideas in your notebook. Using your notes, do all of the **homework** assignments.
5. Do additional **practice midterm exams and other review problems** in the lecture notes, modules (Laulima) and Study Area (Mastering Chemistry).
6. Plan on **spending at least 7 hours per week on this course outside of regular class meeting times**. Here is how your time should be allocated during most weeks:
 - 2-3 hours reading chapter notes, tutorials and DVD text
 - 1-2 hours participating in Supplemental Instruction sessions
 - 3-4 hours doing homework and practice problems

OTHER POLICIES

1. Lecture topics and exam dates are found in the **Course Schedule** at the end of this syllabus.
2. You are expected to have the required **mathematics skills** for the course. You should be familiar with setting up and solving algebraic equations, exponents, logarithms, scientific (engineering) notation, significant figures, proportionality, and percentages. See the math review modules on the course website to review this material. Another place to review is the Study Area on the Mastering Chemistry site. Please let me know immediately if you have any problems with any of these.
3. **Missed Quizzes:** If you are **absent from class**, the quiz you missed will be counted as **zero**. There will be no make-ups for missed quizzes. **Your two lowest quiz scores will be dropped.**
4. **Missed Exams:** If you miss a **midterm exam**, it will count as a **zero**. If a legitimate **emergency** comes up, you must notify me **before the exam** (in person or by email). For a **pre-excused** absence, you may be assigned a score calculated from your other exam scores or allowed to take a make-up exam, provided it is taken before the exams are returned to the other students. It will **not** be possible to make up the **final exam**.

5. **Extra Credit:** You can earn extra credit up to a **maximum of 12 points (2% of total points possible for the course)**. Extra credit can be earned by attending SI sessions beyond the minimum of 8 sessions that are required (0.5 points for each additional session attended). Other opportunities for earning extra credit points (for example, attending a chemistry forum on campus and submitting a written summary of the topic) will be announced during class.
6. You have access to your scores and grades 24/7 in the **gradebook on Laulima**.
7. **Communicating with Instructor:** The best way to reach me is by email and/or by coming to see me during my office hours. Time spent during office hours will be more efficient if you prepare ahead of time and come with specific questions ready to ask.
8. **Disruptive behavior** leads to loss of learning time. It is unfair to the other students in class as well as the instructor and it will not be tolerated. Examples include activated cell phones, checking /sending text messages, chattering/whispering during the course of the lecture, making offensive remarks, eating or drinking in the classroom, packing of books or otherwise making noise, leaving class early, sleeping in class, reading other materials not related to the course, etc. If a student engages in disruptive behavior, the instructor reserves the right to ask the student to leave class immediately and the student will be marked absent.
9. If you have any **special learning needs**, including hearing/visual impairment, please inform the instructor as soon as possible
10. **ZERO TOLERANCE for cheating or academic dishonesty.** See the note regarding academic dishonesty on the following page.
11. Any changes to the Course Schedule will be announced in class at least one week prior to the affected date(s). **You are responsible**, however, for knowing about any changes that are made -- whether or not you were in class for the announcement.

DISABILITIES ACCOMMODATION

If you have a physical, sensory, health, cognitive, or mental health disability that could limit your ability to fully participate in this class, you are encouraged to contact the Disability Specialist Counselor (Ann Lemke) to discuss reasonable accommodations that will help you succeed in this class. She can be reached at 235-7448 or lemke@hawaii.edu. You can also drop by her office in 'Akoakoa 213.

SOME FINAL WORDS OF ADVICE...

BE SURE TO KEEP UP WITH THE WORK IN THIS CLASS! We will be covering a lot of material at a relatively fast pace, so things will become very difficult if you fall behind. Gaining an understanding of basic chemistry concepts and an ability to solve chemistry problems requires practice, and you need to be actively involved in the learning process. This means being focused during the lectures, working through additional practice problems on your own, asking for help when you need it, etc. If you are having trouble keeping up with the class material and wait until the last minute (*i.e.*, right before the exam) before trying to cram everything in, it will be too late.

VERY IMPORTANT NOTE REGARDING ACADEMIC HONESTY

Make sure that you are familiar with the sections related to “Academic Dishonesty” in the College’s policies governing student conduct (available on the WCC website). The fundamental principle governing academic integrity and academic dishonesty is that **each student is responsible for presenting his/her own work at all times.**

It is fine to discuss homework assignments with other students and help each other out – in fact, I strongly encourage you to study with your classmates outside of class time. But it is also important that you learn how to solve problems on your own, and **you must submit your own work.**

Of course it is not OK to collaborate on exams. The following rules will be enforced during exam periods:

- Absolutely no talking once the exam begins. If you have a question or if you need something during an exam, do not ask your neighbor. Raise your hand and I’ll come help you.
- Keep your eyes on your own paper. If I see you looking at someone else’s paper during the quizzes and exams, I will assume you are cheating.
- You are not allowed to bring in any notes or other outside materials to the exams. I will give you copies of the periodic table and other information -- formulas, constant values, etc. (during the lectures, I will tell you which things you need to memorize and which things will be provided for the exams).
- You can (and should) bring a calculator for the exams. But you will only be allowed to use standard scientific calculators – no cell phones, PDA’s (iPhones, Blackberrys, etc.), mini-computers, or any device that can connect to the internet, communicate with other devices, or has data storage capacity.
- No listening to any audio devices (iPods, etc.) during exams.

If you are observed cheating on any of the class assignments (homework, quizzes or exams), your will receive an F for the assignment and I will refer the matter to the Department Head and the Office of the Dean. Cheating is unfair to everyone involved: the teacher, the cheater, and especially the honest students in the class. I adhere to a **zero-tolerance** policy regarding cheating and academic dishonesty, so consider this your first and only warning – there will be no “second chances” in this area.

Trust me – you do NOT want to test me on this!!! I have come down hard on students in my classes for cheating before and will not hesitate to do so if necessary in the future.

WEEK	DATE	TOPICS
1	Jan 13 (Mon)	Ch A: Measurements: Units, Uncertainty, Significant Figures
	Jan 15 (Wed)	Ch A: Sig Figs in Calculations; Units; Dimensional Analysis; Density
2	Jan 20 (Mon)	HOLIDAY – MARTIN LUTHER KING JR DAY
	Jan 22 (Wed)	Ch 1: States of Matter; Elements, Compounds and Mixtures; Atoms & Molecules
3	Jan 27 (Mon)	Ch 1: Atomic Structure; The Periodic Table; Isotopes; Ions
	Jan 29 (Wed)	Ch 1: Charges of Common Ions; Polyatomic Ions; Ionic and Molecular Compounds; Naming and Formulas of Compounds
4	Feb 3 (Mon)	Ch 2: Mass and Moles; Avogadro's Number; Atomic Mass; Molar Mass
	Feb 5 (Wed)	Ch 2: Percent Composition; Empirical and Molecular Formulas; Balancing Chemical Equations
5	Feb 10 (Mon)	Ch 2: Stoichiometry; Theoretical Yield, Limiting Reactant and Percent Yield
	Feb 12 (Wed)	MIDTERM 1
6	Feb 17 (Mon)	HOLIDAY – PRESIDENTS DAY
	Feb 19 (Wed)	Ch 3: Solutions; Concentration and Molarity; Solubility Rules; Ionic and Precipitation Equations; Electrolytes
7	Feb 24 (Mon)	Ch 3: Acids and Bases; Volumetric and Gravimetric Analysis
	Feb 26 (Wed)	Ch 4: Oxidation-Reduction; Oxidation Number
8	Mar 3 (Mon)	Ch 4: Balancing Redox Equations
	Mar 5 (Wed)	Ch 5: Kinetic-Molecular Theory; Gas State Variables; Relationships Between Gas Pressure, Volume, Temperature and Number of Moles
9	Mar 10 (Mon)	Ch 5: Combined Gas Laws; Ideal Gas Law; Gas Density; Gas Stoichiometry
	Mar 12 (Wed)	Ch 5: Kinetic Molecular Theory Revisited; Diffusion and Effusion; Real (Non-Ideal) Gas Behavior
10	Mar 17 (Mon)	Ch 6: Energy, Work; Heat and Temperature; Specific Heat
	Mar 19 (Wed)	MIDTERM 2

11	Mar 24 (Mon)	HOLIDAY – SPRING BREAK
	Mar 26 (Wed)	HOLIDAY – SPRING BREAK
12	Mar 31 (Mon)	Ch 6: Thermodynamics and Thermochemistry; System and Surroundings; First Law of Thermodynamics
	Apr 2 (Wed)	Ch 6: State Functions; Heat, Internal Energy, PV Work; Enthalpy
13	Apr 7 (Mon)	Ch 6: Enthalpy of reaction; Enthalpy Diagrams; Hess' Law; Standard Enthalpy Change; Enthalpy of Formation
	Apr 9 (Wed)	Ch 7: Electromagnetic Radiation; Bohr Model of the Hydrogen Atom; Electron Energy
14	Apr 14 (Mon)	Ch 7: Electronic Structure of the Atom: Quantum mechanics, Orbitals, Quantum Numbers
	Apr 16 (Wed)	Ch 7: Electron Spin; Ground and Excited States; Electron Configurations and Orbital Diagrams
15	Apr 21 (Mon)	Ch 7: Electron Configuration of Ions; Effective Nuclear Charge; Periodic Trends
	Apr 23 (Wed)	Ch 8: Ionic and Covalent Bonds
16	Apr 28 (Mon)	MIDTERM 3
	Apr 30 (Wed)	Ch 8: Polar Covalent Bonds; Electronegativity
17	May 5 (Mon)	Ch 8: Lewis Structures; Formal Charges Ch 9: Molecular Geometry and Hybridization
	May 7 (Wed)	Ch 9: Molecular Orbital Model

**FINAL EXAM: Wednesday, May 14
11:30 am to 1:30 pm**

**REMINDERS: Last day withdraw without a "W" grade: Monday, February 3
Last day to withdraw with a "W" grade: Thursday, March 20**