Chem 151L Elementary Survey of Chemistry Lab
1 credit (CRN 61452)
M 2:00 - 4:45 PM Imiloa 111 & 131

INSTRUCTOR: Dave Reeves
OFFICE: Imiloa 130
E-MAIL: dr6@hawaii.edu
OFFICE HOURS: M 1:00-2:00 PM
TELEPHONE: 236-9116
EFFECTIVE DATE: Fall 2012

WINDWARD COMMUNITY COLLEGE MISSION STATEMENT

Windward Community College offers innovative programs in the arts and sciences
and opportunities to gain knowledge and understanding of Hawai‘i and its unique
heritage. With a special commitment to support the access and educational needs
of Native Hawaiians, we provide O‘ahu’s Ko‘olau region and beyond with liberal
arts, career and lifelong learning in a supportive and challenging environment —
inspiring students to excellence.

CATALOG DESCRIPTION

Experiments introducing laboratory techniques and illustrating chemical principles;
supplemented by films, demonstrations, and problem sessions. (3 hrs. lab.)

Prerequisites: Credit or registration in CHEM 151.
WCC: DY

STUDENT LEARNING OUTCOMES

1. Identify/locate laboratory safety equipment and apply laboratory safety procedures.
2. Assemble apparatus to perform common laboratory techniques to verify basic chemistry
   laws on gases, chemical stoichiometry, preparation of solutions, chemical equilibrium
   and others.
3. Use molecular models and technology to investigate chemistry concepts.
4. Make and record accurate observations, precise measurements and calculations applying
   rules on significant figures.
5. Develop hypotheses, use critical thinking to process results and identify sources of error.
6. Apply and articulate the scientific method by preparing a lab report using the standard
   scientific format.

LEARNING RESOURCES

Required Textbook: Chemistry 151L Laboratory Manual by Colmenares Spring 2011
Course Website:  
http://laulima.hawaii.edu (use UH email account login and password)

Other Requirements:  
Scientific Calculator  
Internet Access

**PURPOSE OF THE LABORATORY COURSE**

The chemistry laboratory allows the student to understand some of the theories discussed in the lecture more thoroughly. In the laboratory you will be involved with the processes of scientific inquiry used to discover chemical principles. It is the only way for the student to learn the techniques that are so important in research and in most laboratories. The student will discover that doing quality work in the laboratory requires a great deal of patience and care.

**GRADING**

The final grade will be based on the following scheme:

<table>
<thead>
<tr>
<th>Component</th>
<th>Percentage</th>
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<tbody>
<tr>
<td>13 Lab reports (including 3 Formal)</td>
<td>70%</td>
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<tr>
<td>2 Long Tests</td>
<td>30%</td>
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<tr>
<td><strong>Total</strong></td>
<td><strong>100%</strong></td>
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</tbody>
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Course grades will be assigned as follows:

A-- 90-100%  B-- 80-89%  C-- 70-79%  D-- 60-69%  F-- less than 60%.

For grades I, W, Cr, NC-- See college catalog

**MODE OF INSTRUCTION**

The primary mode of instruction is through the use of hands-on laboratory activities. Each hands-on activity will be preceded by a class discussion based on a pre-lab homework consisting of a reading assignment and completion of pre-lab questions and problems. During the discussion, videos, websites and computer programs may be utilized.

Conduct of the Lab:

First 30 min  
Discuss answers and review *Pre-Lab* assignment problems and questions, review important principles and *safety* precautions, demonstrate new *techniques* needed or use of new equipment (in Imiloa 111)

Remaining time  
Perform *hands-on* activity, record observation and data, calculate results

10 min  
Clean-up

The activity report is due at the start of the next meeting.
COURSE POLICIES AND TIPS TO IMPROVE LAB SKILLS

1. The Course Content and Schedule (activity titles and corresponding dates) is listed on the last page. Changes in the schedule will be announced in class and in Assignments on Laulima at least a week prior to the affected date. You are responsible for obtaining this information whether or not you were present during the announcement.

2. Before class. Each hands-on activity is described in detail in the Lab Manual (LM). Be prepared for the activity by reading thoroughly the background and pre-lab material on Laulima and take the online quiz. Do the Pre-Lab homework before coming to class. This will allow you to become familiar with the experiment and to better understand the calculations involved. Bring your Lab Manual, scientific calculator and lecture textbook to class at all times.

3. Pre-lab Period (Imiloa 111). You will take your pre-lab quiz on Laulima the day before the lab. The Pre-Lab and quiz constitutes 10% of the report grade. Answers to the Pre-Lab problems and questions will be discussed individually with the instructor at the beginning of the lab period.

4. Important background information about the experiment and safety will be discussed in the Pre-Lab discussion. If the chemistry concept or theory is entirely new to the class, the instructor will give a short discussion. During the Pre-Lab, you are expected to ask questions regarding the procedure and calculation. However, if you have a lot of questions, get help from your study group, a tutor or from the instructor during office/consultation hours prior to coming to class. If you arrive late, ½ a point will be deducted from your lab score.

5. Some experiments require the use of laptop computers from the metal lab cart. This will be wheeled out of the Imiloa AV room before the class and returned there after class.

6. During lab (Imiloa 131). Wear safety goggles and close-toed shoes at all times. Sandals or slippers are not allowed. You will NOT be permitted in the lab if you are not wearing close-toed shoes. Follow laboratory rules and procedures at all times. Treat all chemicals with respect and wipe up any spill, return all reagents to the proper place, replace lids on bottles and report any accident or problem to the instructor. Follow the directions in the Procedure precisely. Don’t take short cuts nor fake results as these are readily spotted.

7. In the lab, you are to work in pairs. Work cooperatively and maintain a positive attitude. Treat it as an opportunity to learn and do not simply rush through an experiment in order to get out of the laboratory as quickly as possible. Maintain a positive attitude and work cooperatively with other students and the laboratory instructor. Be alert and maintain presence of mind.

8. Inappropriate behavior such as activated beepers and cell phones, making offensive remarks, prolonged chattering, etc. will not be tolerated. Do not use cell phones in the classroom or in the laboratory.

9. When recording a measurement, it is expected that the value reflect the precision of the instrument used (never round-off) and affix the correct unit. Record all your data neatly in ink and with units. Do not erase original data. If you make a mistake just put a line thru them. Point deductions will be applied to data with incorrect precision.

10. Do a quick calculation to check if results are reasonable before dismantling the setup. Repeat the experiment if you make a mistake. Consult the instructor to discuss the probable causes of error before doing the repeat.
11. Use laboratory time efficiently and bear in mind that the experiment should be done at least ten minutes before the end of class for cleanup activity. Leave your laboratory station in good order when you are finished for the day. Points will be deducted when safety precautions, chemical transfer, clean-up and waste disposal techniques and procedures are not followed. 

12. After the lab. Complete your calculations as soon as possible after the laboratory. Do not procrastinate and wait until the night before the next laboratory to write your report. Remember you will need to prepare for the new one as well. 

13. Laboratory reports. You are expected to turn in an individual report at the beginning of next class. Lab reports will not be accepted via email. You are expected to discuss results, calculations and interpretations with your laboratory partner and classmates, but calculations and answers in the report should be completely your own work. Copying someone else’s work or answers is cheating. Students involved in cheating will be assigned a grade of “F” and will be reported to the Dean. Hence, the student should not allow anybody to copy one’s work. 

14. Formal Laboratory Reports. A formal laboratory report is required for Experiments on Determining Chemical Formula and Boyle’s Law while an informal lab report is required for all the other experiments. Formal reports should follow the standard format (see sample formal report in Laulima course website). The rubric for grading the formal lab reports and tips on how to write a formal report are found in your Lab Manual. 

15. Late Reports. If submitted one week after the due date, the lab report (only for labs the student was present) will be given a grade of 70% if complete. The grade assigned will be zero if submitted much later. 

16. The graded lab report will be returned to you in the following lab session and you are to return them to the instructor immediately after reviewing them. Graded labs will be returned to you to study a week before each exam. Monitor your class performance by keeping a record of your scores. 

17. Make-up. No make-up is allowed and the student gets “zero” for the missed lab activity. 

18. Attendance will be checked at the beginning of each Pre-Lab session. If you are absent, you will not get a grade for the experiment done on that date. 

19. If you have any special learning needs, including hearing/visual impairment, please inform the instructor as soon as possible. 

**DISABILITIES ACCOMMODATION**

If you have a physical, sensory, health, cognitive, or mental health disability that could limit your ability to fully participate in this class, you are encouraged to contact the Disability Specialist Counselor to discuss reasonable accommodations that will help you succeed in this class. Ann Lemke can be reached at 235-7448, [lemke@hawaii.edu](mailto:lemke@hawaii.edu), or you may stop by Hale ‘Akoakoa 213 for more information.
## COURSE CONTENT AND SCHEDULE

Holidays: Sep 3 (M), Nov 12 (M), Nov 22 (R), Nov 23 (F)
Important Dates: Last day for withdrawal, Oct 29 (M)
Last day of instruction, Dec 6 (R)

<table>
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<tr>
<th>Date</th>
<th>Activity Title</th>
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<tr>
<td>Aug 20</td>
<td>1- Laboratory Safety, Equipment &amp; Procedures</td>
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<td>Aug 27</td>
<td>2- Measurements: Length, Mass, Volume, &amp; Density</td>
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<td>Sept 3</td>
<td>Labor Day Holiday</td>
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<td>Sep 10</td>
<td>3- Endothermic and Exothermic Reactions (Computer-based Lab)</td>
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<td>Sep 17</td>
<td>4- Periodic Trends</td>
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<td>Sep 24</td>
<td>5- Chemical Bonds, Models &amp; Shapes</td>
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<td>Oct 1</td>
<td>6- Types of Chemical Reactions</td>
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<td>Oct 8</td>
<td>7- Determining the Chemical Formula of Magnesium Oxide**</td>
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<td>Oct 15</td>
<td>Midterm Exam (Activities 1-6)</td>
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<td>Oct 22</td>
<td>8- Chemical Stoichiometry</td>
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<td>Oct 29</td>
<td>9- Boyle’s Law** (Computer-based Lab)</td>
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<td>Nov 5</td>
<td>10- Solutions: Preparation and Properties</td>
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<td>Nov 12</td>
<td>Veterans Day Holiday</td>
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<td>Nov 19</td>
<td>11- Acids, Bases, pH Indicators and Buffers</td>
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<td>Nov 26</td>
<td>12- Forensic Detective Activity</td>
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<td>Dec 3</td>
<td>Final Long Exam (Activities 7-12) &amp; Check-out</td>
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*Subject to change
** Require formal lab report